A Coats-Based Formulation for Multiphase Reactive Transport: Application to Hydrogen Storage in Porous Media

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Objectives

- Extend the Coats formulation to multiphase reactive transport,
- Properly handle phase disappearance (especially the aqueous phase),
- Use a Globally Implicit Approach.

Organization

1. Reactive transport model
2. Phase appearance and disappearance
3. Numerical results

Notations (Morel's formalism)

- Primary species:

$$\mathcal{E}^{\mathbf{pr}} = \{e_1^{\mathbf{pr}}, \dots, e_M^{\mathbf{pr}}\}.$$

- Secondary species:

$$\mathcal{E}^{\mathbf{sec}} = \{e_1^{\mathbf{sec}}, \dots, e_{N-M}^{\mathbf{sec}}\},$$

such that each secondary species is associated to the chemical reaction:

$$e^{\mathbf{sec}} \leftrightarrows \sum_{e^{\mathbf{pr}} \in \mathcal{E}^{\mathbf{pr}}} E_{e^{\mathbf{pr}}, e^{\mathbf{sec}}}\, e^{\mathbf{pr}},$$

where

$$E := [E_{e^{\mathbf{pr}}, e}]_{e^{\mathbf{pr}} \in \mathcal{E}^{\mathbf{pr}}, e \in \mathcal{E}} \in \mathbb{R}^{\mathcal{E}^{\mathbf{pr}} \times \mathcal{E}}.$$

- We set $\mathcal{E} := \mathcal{E}^{\mathbf{pr}} \cup \mathcal{E}^{\mathbf{sec}}$.

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- Primary species:
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such that each secondary species is associated to the chemical reaction:
$$e^{\mathbf{sec}} \leftrightarrows \sum_{e^{\mathbf{pr}} \in \mathcal{E}^{\mathbf{pr}}} E_{e^{\mathbf{pr}}, e^{\mathbf{sec}}}\, e^{\mathbf{pr}},$$
where
$$E := [E_{e^{\mathbf{pr}}, e}]_{e^{\mathbf{pr}} \in \mathcal{E}^{\mathbf{pr}}, e \in \mathcal{E}} \in \mathbb{R}^{\mathcal{E}^{\mathbf{pr}} \times \mathcal{E}}.$$
- We set $\mathcal{E} := \mathcal{E}^{\mathbf{pr}} \cup \mathcal{E}^{\mathbf{sec}}$.
- To link a species $e \in \mathcal{E}$ to its phase $\beta \in \mathcal{P}$:
$$\alpha : e \in \mathcal{E} \mapsto \alpha(e) \in \mathcal{P}.$$

Total conservation equation

Total conservation

$$\sum_{e\in\mathcal{E}} E_{e^{\mathrm{pr}},e}\partial_t n_e + \sum_{e\in\mathcal{E}^{\mathrm{f}}} E_{e^{\mathrm{pr}},e}\,\mathrm{div}\left(\zeta^{\alpha(e)} C_e \mathrm{u}^{\alpha(e)}\right) = 0, \quad \forall e^{\mathrm{pr}} \in \mathcal{E}^{\mathrm{pr}},$$

where

- $\mathcal{E}^{\mathrm{f}}$: set of species in fluid phases,
- n_e $[\mathrm{mol}\cdot\mathrm{m}^{-3}]$: volumetric molar concentration of e, equals to

$$n_e = \phi\zeta^{\alpha(e)} S^{\alpha(e)} C_e, \quad \forall e \in \mathcal{E}^{\mathrm{f}},$$

- ϕ $[-]$: porosity,
- ζ^β $[\mathrm{mol}\cdot\mathrm{m}^{-3}]$: phase molar density,
- S^β $[-]$: phase saturation,
- C_e $[-]$: mole fraction of e in its phase,
- u^β $[\mathrm{m}\cdot\mathrm{s}^{-1}]$: follows the Darcy's law.

Chemical reactions

- Phase transition: Fugacity equality.
- Chemical equilibrium: Law of mass action,

$$\frac{1}{a_{e^{\mathbf{sec}}}} \prod_{e^{\mathbf{pr}} \in \mathcal{E}^{\mathbf{pr}}} \left(a_{e^{\mathbf{pr}}}\right)^{E_{e^{\mathbf{pr}},R}} = K_R.$$

- Kinetic reaction: Conservation equation with reaction rate T_e,

$$\partial_t n_e + \operatorname{div}\left(\zeta^{\alpha(e)} C_e \mathbf{u}^{\alpha(e)}\right) = T_e, \text{ if e is a fluid kinetic species,}$$
$$\partial_t n_e = T_e, \text{ if e is a solid kinetic species.}$$

Closure equations

- Total volume conservation:

$$\underbrace{\phi}_{\text{fluid}} + \underbrace{\phi^{\mathbf{s}} + 1 - \phi_{\mathbf{T}}}_{\text{solid}} = 1 \quad \Leftrightarrow \quad \phi + \phi^{\mathbf{s}} = \phi_{\mathbf{T}},$$

where

  - $\phi^{\mathbf{s}}$: volume fraction of reactive rock,
  - $1 - \phi_{\mathbf{T}}$: volume fraction of non-reactive rock.

- Closure relations:

$$\sum_{\beta \in \mathcal{P}^{\mathbf{f}}} S^{\beta} = 1, \qquad \sum_{e \in \alpha^{-1}(\{\beta\})} C_e = 1, \ \forall \beta \in \mathcal{P}^{\mathbf{f}} \qquad \text{and} \qquad \phi^{\mathbf{s}} = \sum_{m \in \mathcal{P}^{\mathbf{s}}} \frac{n_m}{\zeta^m}.$$

Summary

- $\#\mathcal{E} + \#\mathcal{P}^{\mathbf{f}} + 2$ equations:
  - $\#\mathcal{E}^{\mathbf{pr}}$: the conservation equations for total primary species,
  - $\#\mathcal{E}^{\mathbf{sec}}$: one algebraic or differential equation for each equilibrium or kinetic chemical reaction, respectively,
  - $\#\mathcal{P}^{\mathbf{f}} + 2$: closure equations,

- $\#\mathcal{E} + \#\mathcal{P}^{\mathbf{f}} + 2$ natural variables (Coats formulation):
  - pressure P,
  - saturations $S^{\beta}, \quad \beta \in \mathcal{P}^{\mathbf{f}}$,
  - solid porosity $\phi^{\mathbf{s}}$,
  - molar fractions $C_e, \quad e \in \mathcal{E}^{\mathbf{f}}$,
  - molar concentrations $n_e, \quad e \in \mathcal{E}^{\mathbf{s}}$.

Summary

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  - saturations $S^{\beta}, \quad \beta \in \mathcal{P}^{\mathbf{f}}$,
  - solid porosity $\phi^{\mathbf{s}}$,
  - molar fractions $C_e, \quad e \in \mathcal{E}^{\mathbf{f}}$,
  - molar concentrations $n_e, \quad e \in \mathcal{E}^{\mathbf{s}}$.

Remark

Here, we assume that all phases are present!

Organization

1. Reactive transport model
2. Phase appearance and disappearance
3. Numerical results

Discretization

- Finite volume with TPFA and upwind scheme.

Discretized conservations of total primary species

For $e \in \mathcal{E}^{\mathsf{pr}}$,

$$\sum_{e' \in \mathcal{E}} E_{e,e'}|K| \left(\chi_K(e')\, n_{e',K} - \chi_K^{n-1}(e')\, n_{e',K}^{n-1} \right) + \Delta t^n \sum_{e' \in \mathcal{E}^{\mathsf{f}}} \sum_{\sigma = K|L \in \mathcal{F}_K \cap \mathcal{F}_h} \chi_{K_\sigma^{\alpha(e')}}(e') E_{e,e'} \left(\zeta^{\alpha(e')} C_{e'} \eta^{\alpha(e')} \right)_{K_\sigma^{\alpha(e')}} V_{KL}^{\alpha(e')} = 0,$$

where

- $\mathcal{Q}_K = (\mathcal{Q}_K^{\mathsf{f}}, \mathcal{Q}_K^{\mathsf{s}})$: set of present fluid/solid phases in K,
- $\chi_K(e) := 1_{\mathcal{Q}_K}(\alpha(e))$ (analog for $\chi_K^{n-1}(e)$).

Discretization

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- $\chi_K(e) := 1_{\mathcal{Q}_K}(\alpha(e))$ (analog for $\chi_K^{n-1}(e)$).

Problem

Singularity issue if a primary species disappears.

Discretization

- Additional concentration unknowns chosen based on the rank deficiency of the formula matrix restricted to the present species:

$$\widetilde{n}_{e,K}, \quad e \in \widetilde{\mathcal{E}}_{\mathcal{Q}_K}.$$

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Discretized conservations of total primary species

For $e \in \mathcal{E}^{\mathrm{pr}}$,

$$1_{\widetilde{\mathcal{E}}_{\mathcal{Q}_K}}(e)|K|\ \widetilde{n}_{e,K} + \sum_{e' \in \mathcal{E}} E_{e,e'}|K| \left(\chi_K(e')\, n_{e',K} - \chi_K^{n-1}(e')\, n_{e',K}^{n-1}\right)$$
$$+ \Delta t^n \sum_{e' \in \mathcal{E}^{\mathrm{f}}} \sum_{\sigma = K|L \in \mathcal{F}_K \cap \mathcal{F}_h} \chi_{K_\sigma^{\alpha(e')}}(e') E_{e,e'} \left(\zeta^{\alpha(e')} C_{e'} \eta^{\alpha(e')}\right)_{K_\sigma^{\alpha(e')}} V_{KL}^{\alpha(e')} = 0.$$

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$$1_{\widetilde{\mathcal{E}}_{\mathcal{Q}_K}}(e)|K|\ \widetilde{n}_{e,K} + \sum_{e' \in \mathcal{E}} E_{e,e'}|K| \left(\chi_K(e')\, n_{e',K} - \chi_K^{n-1}(e')\, n_{e',K}^{n-1}\right)$$
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Remarks

- Regularize the Jacobian matrix,
- Track incoming flux from surrounding cells (phases reappearance).

Phase appearance and disappearance

Phase removal

- Fluid phase β disappears if $S^\beta < 0$.
- Mineral phase m disappears if $n_m < 0$.

Phase appearance and disappearance

Phase removal

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Phase appearance with liquid phase

- A mineral phase appears when its solubility product exceeds equilibrium.
- A gas phase appears if

$$\widetilde{n}_e > 0$$

for any gas species e in $\widetilde{\mathcal{E}}_{\mathcal{Q}_K}$, or if

$$\sum_e \widetilde{C}_e > 1.$$

Phase appearance and disappearance

Phase removal

- Fluid phase β disappears if $S^\beta < 0$.
- Mineral phase m disappears if $n_m < 0$.

Phase appearance with liquid phase

- A mineral phase appears when its solubility product exceeds equilibrium.
- A gas phase appears if

$$\widetilde{n}_e > 0$$

for any gas species e in $\widetilde{\mathcal{E}}_{\mathcal{Q}_K}$, or if

$$\sum_e \widetilde{C}_e > 1.$$

Without liquid phase

A multiphase flash calculation is performed at fixed total composition.

Organization

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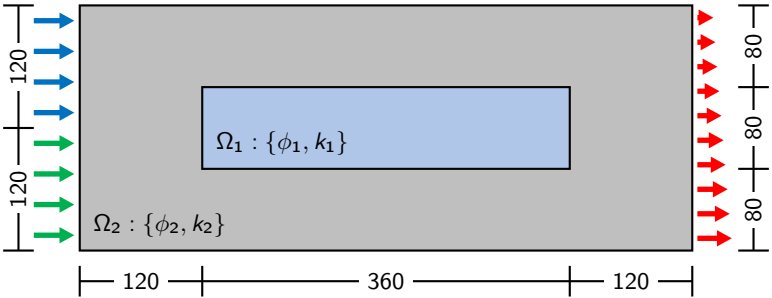
Settings (CO_2 test case)



Gas stream (pure CO_2^g):	Q_{gas} [m^3/day]
Water stream (pure H_2O^ℓ):	Q_{wat} [m^3/day]
Outflow (hydrostatic pressure):	p_{out} [bar]

Settings (CO_2 test case)

- Primary species (4):

$$\mathcal{E}^{\mathrm{pr}} = \{H_2O^\ell, CO_2^\ell, H^+, Ca^{2+}\},$$

- Reactions (6):

$$\begin{aligned}
HCO_3^- &\leftrightarrows H_2O^\ell + CO_2^\ell - H^+, && \text{(equilibrium aq-aq)}\\
OH^- &\leftrightarrows H_2O^\ell - H^+, && \text{(equilibrium aq-aq)}\\
CO_3^{2-} &\leftrightarrows H_2O^\ell + CO_2^\ell - 2H^+, && \text{(equilibrium aq-aq)}\\
H_2O^g &\leftrightarrows H_2O^\ell, && \text{(phase change)}\\
CO_2^g &\leftrightarrows CO_2^\ell, && \text{(phase change)}\\
CaCO_3 &\leftrightarrows Ca^{2+} + H_2O^\ell + CO_2^\ell - 2H^+, && \text{(kinetic min-aq)}
\end{aligned}$$

- Phases: $\mathcal{P} = \{\ell, g, m_{CaCO_3}\}$,

Additional unknowns (CO_2 test case)

$$E := \left(\begin{array}{cccc|cccccc} H_2O^\ell & CO_2^\ell & H^+ & Ca^{2+} & HCO_3^- & OH^- & CO_3^{2-} & H_2O^g & CO_2^g & CaCO_3 & \\ 1 & 0 & 0 & 0 & 1 & 1 & 1 & 1 & 0 & 1 & H_2O^\ell \\ 0 & 1 & 0 & 0 & 1 & 0 & 1 & 0 & 1 & 1 & CO_2^\ell \\ 0 & 0 & 1 & 0 & -1 & -1 & -2 & 0 & 0 & -2 & H^+ \\ 0 & 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 1 & Ca^{2+} \end{array}\right).$$

Additional unknowns (CO_2 test case)

$$E := \begin{array}{c} \begin{array}{cccc|cccccc} H_2O^\ell & CO_2^\ell & H^+ & Ca^{2+} & HCO_3^- & OH^- & CO_3^{2-} & H_2O^g & CO_2^g & CaCO_3 \end{array} \\ \left(\begin{array}{cccc|cccccc} 1 & 0 & 0 & 0 & 1 & 1 & 1 & 1 & 0 & 1 \\ 0 & 1 & 0 & 0 & 1 & 0 & 1 & 0 & 1 & 1 \\ 0 & 0 & 1 & 0 & -1 & -1 & -2 & 0 & 0 & -2 \\ 0 & 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 1 \end{array}\right) \end{array} \begin{array}{c} H_2O^\ell \\ CO_2^\ell \\ H^+ \\ Ca^{2+} \end{array}.$$

$$E_{\mathcal{Q}_K} = \begin{array}{c} \begin{array}{cc} H_2O^g & CO_2^g \end{array} \\ \left(\begin{array}{cc} 1 & 0 \\ 0 & 1 \\ 0 & 0 \\ 0 & 0 \end{array}\right) \end{array}, \text{ if } \mathcal{Q}_K = \{g\}.$$

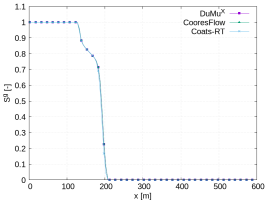
$$E_{\mathcal{Q}_K} = \begin{array}{c} \begin{array}{ccc} H_2O^g & CO_2^g & CaCO_3 \end{array} \\ \left(\begin{array}{ccc} 1 & 0 & 1 \\ 0 & 1 & 1 \\ 0 & 0 & -2 \\ 0 & 0 & 1 \end{array}\right) \end{array}, \text{ if } \mathcal{Q}_K = \{g, m_{CaCO_3}\}$$

- Rank: 2,
- Two additional unknowns: $\widetilde{n}_{H^+}$ and $\widetilde{n}_{Ca^{2+}}$.

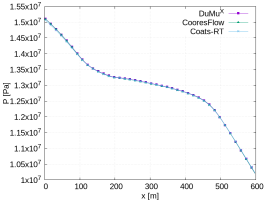
- Rank: 3,
- One additional unknown: $\widetilde{n}_{H^+}$ (or $\widetilde{n}_{Ca^{2+}}$).

Results

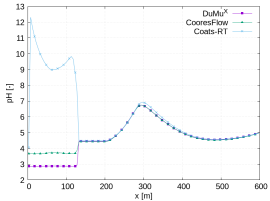
Code team	TS (failed)	NI	LI / NI	F (failed)	CPU	Architecture
DuMu^X	17571 (2937)	3.45	8.37	-	66124	Intel i7-8565U @1.8 GHz
CooresFlow	2008 (0)	4.1	not available	-	10224	Intel Xeon E5-1620 v3 @3.5 GHz
Coats-RT	2032 (8)	5.2	11.76	206939 (0)	6361	Intel Core Ultra 7 255H @2.0 GHz



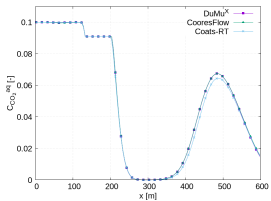
(d) Gaz Saturation



(e) Pressure



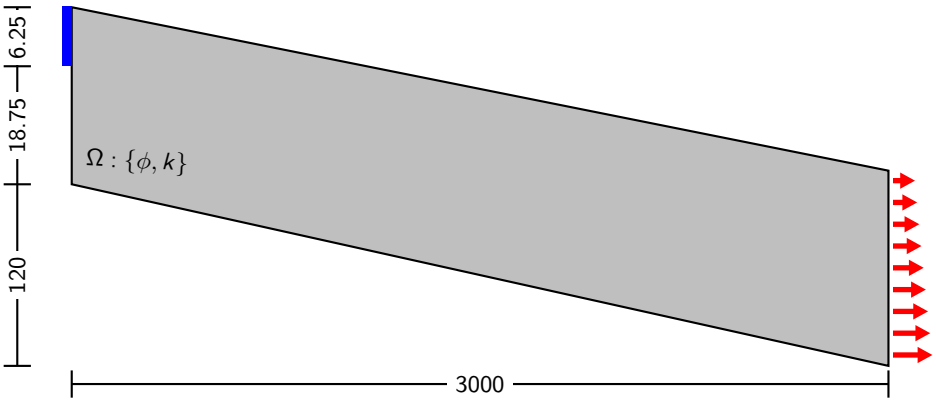
(f) pH



(g) $C_{CO_2^g}$

Figure: Benchmark – DuMu^X, CooresFlow and Coats-RT.

Settings (H_2 test case)



Settings (H_2 test case)

- Primary species (7):

$$\mathcal{E}^{\text{pr}} := \{H_2S^\ell, H_2^\ell, CO_2^\ell, H_2O, H^+, Ca^{2+}, Na^+\}.$$

- Reactions (9):

$$\begin{aligned}
SO_4^{2-} &\leftrightarrows H_2S^\ell - 4H_2^\ell + 4H_2O - 2H^+, && \text{(kinetic aq-aq)}\\
HS^- &\leftrightarrows H_2S^\ell - H^+, && \text{(equilibrium aq-aq)}\\
OH^- &\leftrightarrows H_2O - H^+, && \text{(equilibrium aq-aq)}\\
HCO_3^- &\leftrightarrows CO_2^\ell - H^+ + H_2O, && \text{(equilibrium aq-aq)}\\
CO_3^{2-} &\leftrightarrows CO_2^\ell - 2H^+ + H_2O, && \text{(equilibrium aq-aq)}\\
H_2S^g &\leftrightarrows H_2S^\ell, && \text{(phase change)}\\
H_2^g &\leftrightarrows H_2^\ell, && \text{(phase change)}\\
CO_2^g &\leftrightarrows CO_2^\ell, && \text{(phase change)}\\
CaCO_3 &\leftrightarrows Ca^{2+} + CO_2^\ell - 2H^+ + H_2O, && \text{(equilibrium min-aq)}
\end{aligned}$$

- Phases: $\mathcal{P} = \{\ell, g, m_{CaCO_3}\}$.

Settings (H_2 test case)

$$SO_4^{2-} \leftrightarrows H_2S^\ell - 4H_2^\ell + 4H_2O - 2H^+$$

- Monod-type law: (bacterial activity)

$$T_{SO_4^{2-}} = -v_{max} \frac{10^{-3}\zeta^\ell C_{SO_4^{2-}}}{10^{-3}\zeta^\ell C_{SO_4^{2-}} + K_{SO_4^{2-}}} \frac{10^{-3}\zeta^\ell C_{H_2^\ell}}{10^{-3}\zeta^\ell C_{H_2^\ell} + K_{H_2^\ell}},$$

with

- $v_{max} = 1.45 \times 10^{-5}\ \mathrm{mol \cdot L^{-1} \cdot year^{-1}}$,
- $K_{SO_4^{2-}} = 10^{-4}\ \mathrm{mol \cdot L^{-1}}$, and $K_{H_2^\ell} = \varepsilon\ \mathrm{mol \cdot L^{-1}}$.

Additional unknowns (H_2 test case)

$$E := \left(\begin{array}{ccccccc|ccccccccc} 1 & 0 & 0 & 0 & 0 & 0 & 0 & 1 & 1 & 0 & 0 & 0 & 1 & 0 & 0 & 0 \\ 0 & 1 & 0 & 0 & 0 & 0 & 0 & -4 & 0 & 0 & 0 & 0 & 0 & 1 & 0 & 0 \\ 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 1 & 1 & 0 & 0 & 1 & 1 \\ 0 & 0 & 0 & 1 & 0 & 0 & 0 & 4 & 0 & 1 & 1 & 1 & 0 & 0 & 0 & 1 \\ 0 & 0 & 0 & 0 & 1 & 0 & 0 & -2 & -1 & -1 & -1 & -2 & 0 & 0 & 0 & -2 \\ 0 & 0 & 0 & 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 1 \\ 0 & 0 & 0 & 0 & 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 \end{array}\right).$$

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- If $\mathcal{Q}_K = \{g\}$:

$$E_{\mathcal{Q}_K} = \begin{array}{c} \begin{array}{ccc} H_2S^g & H_2^g & CO_2^g \end{array} \\ \left(\begin{array}{ccc} 1 & 0 & 0 \\ 0 & 1 & 0 \\ 0 & 0 & 1 \\ 0 & 0 & 0 \\ 0 & 0 & 0 \\ 0 & 0 & 0 \\ 0 & 0 & 0 \end{array}\right) \end{array}.$$

- Rank: 3,
- Four additional unknowns: $\widetilde{n}_{H_2O}$, $\widetilde{n}_{H^+}$, $\widetilde{n}_{Ca^{2+}}$ and $\widetilde{n}_{Na^+}$.

Additional unknowns (H_2 test case)

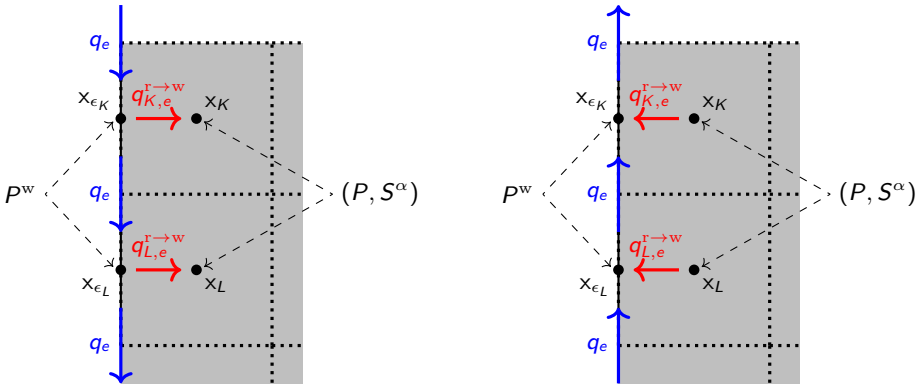
$$E := \left(\begin{array}{ccccccc|ccccccccc} 1 & 0 & 0 & 0 & 0 & 0 & 0 & 1 & 1 & 0 & 0 & 0 & 1 & 0 & 0 & 0 \\ 0 & 1 & 0 & 0 & 0 & 0 & 0 & -4 & 0 & 0 & 0 & 0 & 0 & 1 & 0 & 0 \\ 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 1 & 1 & 0 & 0 & 1 & 1 \\ 0 & 0 & 0 & 1 & 0 & 0 & 0 & 4 & 0 & 1 & 1 & 1 & 0 & 0 & 0 & 1 \\ 0 & 0 & 0 & 0 & 1 & 0 & 0 & -2 & -1 & -1 & -1 & -2 & 0 & 0 & 0 & -2 \\ 0 & 0 & 0 & 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 1 \\ 0 & 0 & 0 & 0 & 0 & 0 & 1 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 & 0 \end{array}\right).$$

- If $\mathcal{Q}_K = \{g, m_{\mathrm{CaCO_3}}\}$:

$$E_{\mathcal{Q}_K} = \begin{array}{c} \begin{array}{cccc} H_2S^g & H_2^g & CO_2^g & CaCO_3 \end{array} \\ \left(\begin{array}{cccc} 1 & 0 & 0 & 0 \\ 0 & 1 & 0 & 0 \\ 0 & 0 & 1 & 1 \\ 0 & 0 & 0 & 1 \\ 0 & 0 & 0 & -2 \\ 0 & 0 & 0 & 1 \\ 0 & 0 & 0 & 0 \end{array}\right) \end{array}.$$

- Rank: 4,
- Three additional unknowns: for example $\widetilde{n}_{\mathrm{H_2O}}$, $\widetilde{n}_{\mathrm{H^+}}$ and $\widetilde{n}_{\mathrm{Na^+}}$.

Simplified vertical well model



(a) Injection (b) Production

Figure: Discretization of the well on a cartesian mesh.

Simplified vertical well model

- Injection:

$$q_{K,e}^{\mathsf{r}\to\mathsf{w}} = \begin{cases} \zeta^g(P_K^{\mathsf{w}})\,\eta^g(1)\,\mathrm{WI}\,\min\left(0, P_K - P_K^{\mathsf{w}}\right) & \text{if } e = \mathrm{H}_2^g, \\ 0 & \text{otherwise,} \end{cases} \leq 0,$$

- Production:

$$q_{K,e}^{\mathsf{r}\to\mathsf{w}} := \zeta^{\alpha(e)}(P_K)\,\eta^{\alpha(e)}\left(S_K^{\alpha(e)}\right) C_{e,K}\,\mathrm{WI}\,\max\left(0, P_K - P_K^{\mathsf{w}}\right) \geq 0,$$

with WI is the well index $[\mathrm{m}]$.

Simplified vertical well model

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with WI is the well index $[\mathrm{m}]$.

Well equation (for injection, analog for production)

$$\min\left(\sum_{e^{\mathsf{pr}}\in\mathcal{E}^{\mathsf{pr}}\cap\mathcal{E}^{\mathsf{f}}} \sum_{K\in\mathcal{M}^{\mathrm{w}}} \sum_{e\in\mathcal{E}^{\mathsf{f}}} E_{e^{\mathsf{pr}},e}\, q_{K,e}^{\mathsf{r}\to\mathsf{w}}\left(P_{\mathrm{bh}}^{\mathsf{w}} + \Delta P_K^{\mathsf{w},n-1}\right) - \overline{Q}, \overline{P} - P_{\mathrm{bh}}^{\mathrm{w}} \right) = 0.$$

where $\mathcal{M}^{\mathrm{w}}$ is the set of reservoir cells connected to the well.

Simplified vertical well model – Well Index

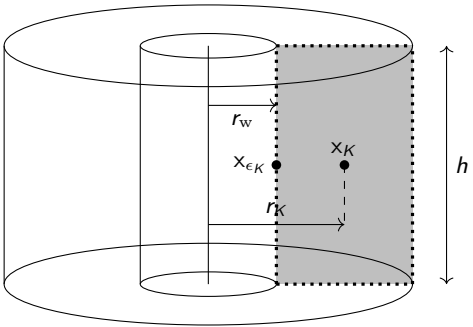
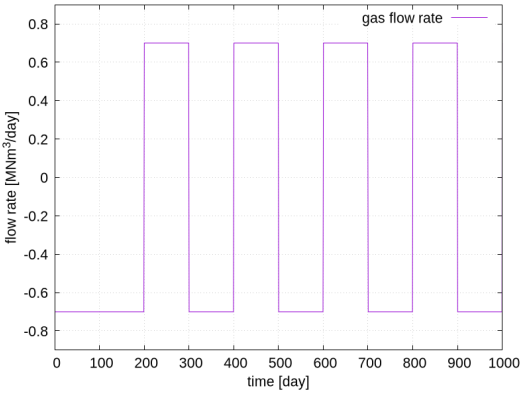


Figure: Discretization in radial geometry.

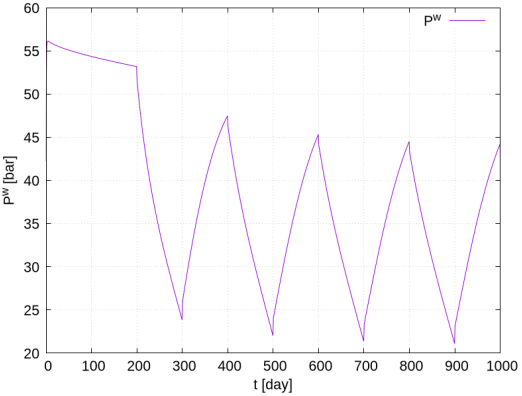
$$\mathrm{WI} = \frac{2\pi h}{\ln\left(r_K / r_w\right)},$$

assuming a homogeneous and isotropic medium (i.e. $\mathbb{K} = k\,\mathrm{Id}$, with k the permeability $[\mathrm{m}^2]$),

Results

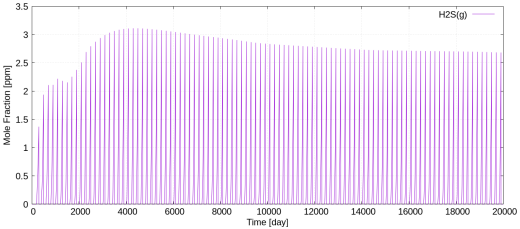


(a) Gaz flow rate

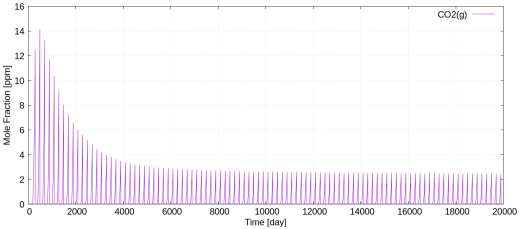


(b) $P^{\mathrm{w}}_{\mathrm{bh}}$

Results

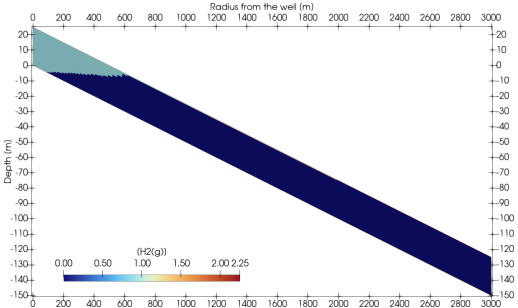


(c) H_2S^g

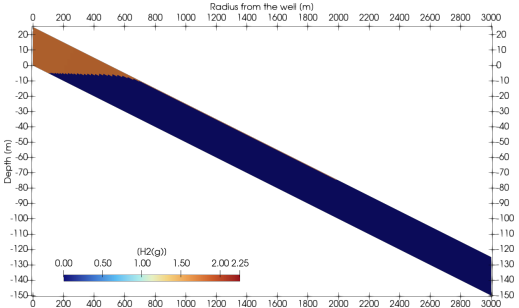


(d) CO_2^g

Results – H_2^g

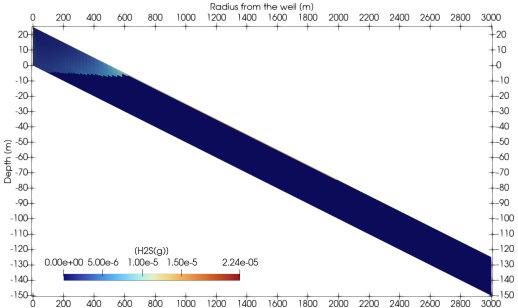


(e) $t = 900$ days

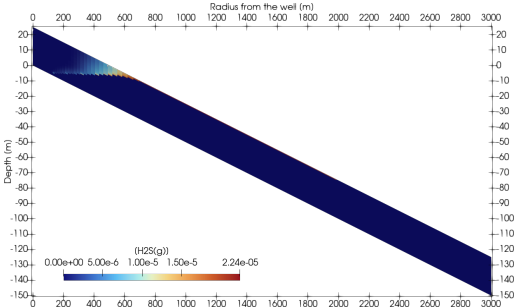


(f) $t = 1000$ days

Results – H_2S^g

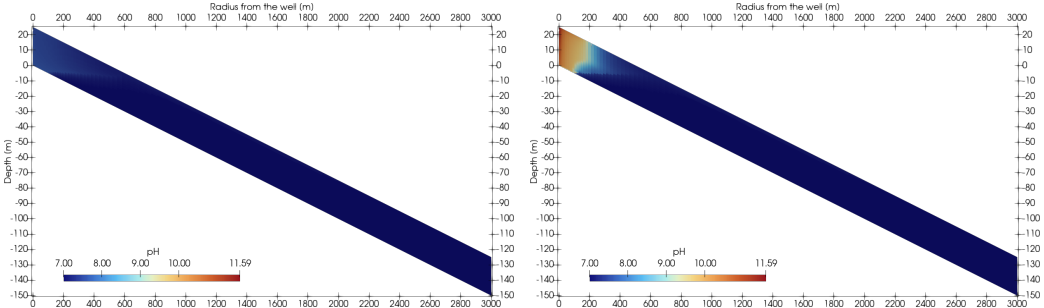


(g) $t = 900$ days



(h) $t = 1000$ days

Results – pH



(i) $t = 900$ days (j) $t = 1000$ days

Conclusion and perspectives

Conclusion

- Fully implicit multiphase reactive transport model based on Morel's and Coats' formulations.
- Robust handling of phase appearance/disappearance, including the aqueous phase.
- Validate the method and the implementation on the CO_2 benchmark.
- Application to a 2D axisymmetric hydrogen storage with injection/production cycles.

Future work

- More realistic H_2 test case (pH buffering reactions, methanation, gazeous H_2O, ...).

Thank you for your attention!